Reg.No. :

I.Answer all the questions.

II.Use Blue pen only.

III.Question No 15,16 is compulsory

Time : 01:00:00 Hrs

Total Marks : 50

Section-A

5 x 1 = 5

1) One can draw the map of building on a glass plate by

    (a) HI   (b) HF   (c) HBr   (d) HCl

2) Among the halogen acid, the weakest acid is

    (a) HF   (b) HCl   (c) HBr   (d) HI

3) Halogens belong to the group number

    (a) 14   (b) 15   (c) 17   (d) 18

4) The noble gases are unreactive because they

    (a) have same number of electrons   (b) have an atomicity of one   (c) are gases with low densities   (d) have stable electronic configuration

5) The shape of XeF_4 is

    (a) tetrahedral   (b) octahedral   (c) square planar   (d) pyramidal

Section-B

6 x 3 = 18

6) Fluorine atom is more electonegative than iodine atom yet, HF has lower acid strength than HI. Explain?

7) What are interhalogen compounds?Give one preparation.

8) Discuss the oxidising power of fluorine.

9) How is XeF_6 prepared?

10) Why do noble gases form compounds with fluorine and oxygen only?

11) Write the uses of Neon?

Section-C

5 x 5 = 25

12) What are silicones? How are they prepared? Mention its uses?

13) Give an account of manufacture of lead. Describe its important properties and uses.

14) How are the following compounds prepared from phosphorus?a) P_2O_3 b) P_2O_5 c) PCl_3 d) PCl_5 e) PH_3

15) a)   Describe in detail how noble gases are isolated from air?

(OR)

    b)   Give an account of various types of compounds which are formed by xenon.

16) a)   How fluorine is isolated from their fluorides?Describe its important properties and uses.

(OR)

    b)   Give a detailed account of the interhalogen compounds with special reference to the compounds involving iodine.Draw their structures.

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